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# Synthesis of $\text{TiO}_2$ -Ag Nanocomposites via Green Route for Dye Degradation and Antimicrobial Studies

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**Abstract:** Synthesis of nanocomposites through green synthesis offers a few advantages over the chemical and physical procedures as it is an easy, fast and eco-friendly alternative that does not involve hazardous chemicals. In this work, we report the green synthesis of  $\text{TiO}_2$ -Ag nanocomposites using titanium tetra isopropoxide (TTIP) and silver nitrate ( $\text{AgNO}_3$ ) as precursors, Aloe vera plant extract as reducing agent. The characterization of nanocomposites is done by X-ray powder diffraction (XRD), UV-Visible spectroscopy, Scanning electron microscopy (SEM) and Fourier transform infrared spectroscopy (FTIR). Electrocatalytic activity of these nanocomposites was examined by degradation of crystal violet dye using potentiostat. The antimicrobial activity was tested against both gram positive and gram negative bacteria viz. *Bacillus subtilis* and *Escherichia coli* respectively. Green synthesized  $\text{TiO}_2$ -Ag nanocomposites were of anatase phase of  $\text{TiO}_2$  and average crystallite size was  $7.06 \pm 3.2$  nm. Complete decolorization of crystal violet dye was achieved in 8 h of electrolysis using  $\text{TiO}_2$ -Ag nanocomposites. Further antimicrobial activity of  $\text{TiO}_2$ -Ag nanocomposites is studied. The diameter of inhibition zone for gram positive *Bacillus subtilis* and gram negative *Escherichia coli* is 13 mm and 21 mm respectively. The nanocomposites exhibited greater antimicrobial activity against the gram negative bacteria *Escherichia coli*. The green  $\text{TiO}_2$ -Ag nanocomposites showed effective electro-catalytic and antimicrobial activity by degrading the dye and inhibiting the bacterial growth respectively.

**Key words:**  $\text{TiO}_2$ -Ag, nanocomposites, green synthesis, dye degradation, antimicrobial, gram positive, gram negative

## I. INTRODUCTION

The textile industries have a high environmental impact, due to the release of large volumes of wastewater that contain high organic content and strong color adding agents. To avoid the dangerous accumulation of dyes in the aquatic environment, present research efforts to develop nanoparticles for their removal in industrial wastewaters [1]. Generally titanium dioxide nanoparticles are used for dye degradation. Titanium dioxide nanoparticles have hydrophobicity, non-toxic, catalyst and large band gap properties [2]. Since ancient times silver has been known for improving immunity in human beings. Silver nanoparticles have attracted great attention due to their unusual optical, chemical, magnetic, electronic, catalyst and antibacterial properties [3]. The noble metal nanoparticles have good dispersity on the particles of the  $\text{TiO}_2$  support, resulting in excellent catalytic activities [4]. Further  $\text{TiO}_2$  support enhances the antibacterial effect of silver nanoparticles. In the present study plant extract is used for synthesizing  $\text{TiO}_2$ -Ag nanocomposites to enhance both catalytic properties (principally for dye degradation) and anti microbial properties. Green synthesis is eco friendly; the product has fewer side effects when used for biomedical uses, low cost and can be used for various applications.

## II. EXPERIMENT

### A. Materials

Fresh and healthy Aloe Vera (*Aloe barbadensis* mill) plants were collected from the garden in Department of Chemical Engineering, JNTUA College of Engineering, Ananthapuramu, Andhra Pradesh. Titanium tetra isopropoxide (TTIP), Silver nitrate ( $\text{AgNO}_3$ ), Sodium sulphate ( $\text{Na}_2\text{SO}_4$ ), and Crystal violet dye were purchased from Sigma Aldrich Chemicals (India) Ltd and are AR grade. Agar-Agar and Peptone were purchased from S D Fine-Chem Limited (India), Sodium chloride ( $\text{NaCl}$ , 99% pure) and Beef extract were purchased from Molychem (India). Nafion (10%) liquid solution is purchased from Dupont Chemicals, USA. Double distilled water was used throughout the experiments. All chemicals were used without any further purification.

### B. Characterizations

X-ray diffraction (XRD) analysis of  $\text{TiO}_2$ -Ag powder was carried out by a WRD PANalytical "XPERT 3 PRO", employing  $\text{Cu-K}\alpha$  radiation of wavelength 0.154 Å. Scanning electron microscopy (SEM) was performed at 15 kV. A Systronics PC-based double

beam UV-Vis spectrophotometer (type: 2201) was used to generate the spectrum for the scan range 190-1000 nm. Fourier Transform Infrared Spectroscopy (FTIR) was conducted to identify the functional groups.

### C. Synthesis of $\text{TiO}_2$ -Ag Nanocomposites

Precursor solution was prepared by continuous stirring of 5 ml of TTIP and 50 ml of ethanol for 1 hr. To this 1 mM  $\text{AgNO}_3$  solution with 30 ml of *Aloe vera* extract was added dropwise under continuous stirring for 1 h. The solution obtained was dried initially at 110 °C for 1 h and later at 200 °C for the same time period, followed by calcination at 500 °C for 1 h. Finally, particles obtained were ground to fine powder by using a mortar and pestle [5].

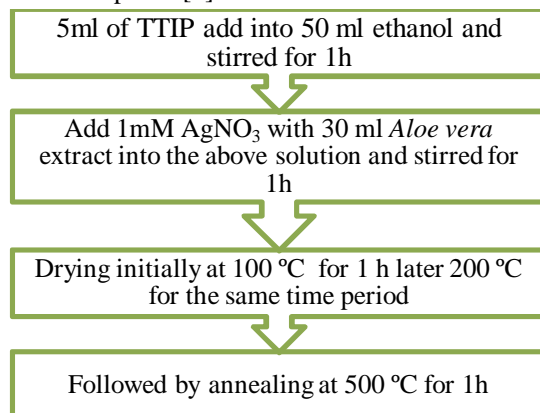


Fig 1: Flow chat of  $\text{TiO}_2$ -Ag nanocomposites synthesis

### D. Electrochemical Analysis

Cyclic voltammetry (CV) experiments were carried out at room temperature using Potentiostat (K-lyte 1.0). The cell used for voltammetric experiments was the three-electrode type (fig 2) Ag/AgCl as a reference electrode, Pt mesh as a counter electrode, and bare Ti and  $\text{Ti/TiO}_2$ -Ag as working electrodes. 55mg/L of crystal violet dye is chosen as electrolyte solution along with 1M  $\text{Na}_2\text{SO}_4$  as supporting electrolyte. Chronoamperometric studies were performed at different fixed voltages at a scan rate of 0.05 V/s. The fixed voltages were obtained from initial CV graph, where the first oxidation peak rises during the forward scan at 0.6 V and at 1.1 V for bare Ti and  $\text{Ti/TiO}_2$ -Ag respectively. UV-Vis spectrophotometer was used to measure the decolorization efficiency at different time intervals and is calculated by using the formula. % decolorisation =  $(A_0 - A) / A_0 \times 100$ , where  $A_0$  and A are the absorbances (at the maximum wavelength,  $\lambda_{\text{max}}$ ) of dye solution before/initial (at  $t = 0$ ) and after electrochemical treatment (at  $t > 0$ ) [6].



Fig 2: Potentiostat



### E. Antimicrobial Studies

*Escherichia coli* (*E.coli*) and *Bacillus subtilis* (*B. subtilis*) bacteria were cultivated in Luria Bertani (LB) medium. First LB medium was prepared by using three compounds that are  $5 \text{ g L}^{-1}$  NaCl,  $5 \text{ g L}^{-1}$  peptone, and  $1.5 \text{ g L}^{-1}$  beef extract (fig 3). For 100 ml of LB medium, 0.5g of NaCl, 0.5 g of peptone, and 0.15 g of beef extract were dissolved in 100 ml distilled water then the solution was autoclaved at  $121^\circ\text{C}$  for 20 min. After cooling the solution was inoculated. To cultivate the bacteria, solution was put into the orbital shaker at  $38^\circ\text{C}$  (room temperature) for 8 h. After orbital shaking solution density increased, this indicated the bacterial growth in the LB medium. This LB medium was stored in refrigerator at  $4^\circ\text{C}$  for further use. Nutrient medium was prepared using  $5 \text{ g L}^{-1}$  NaCl,  $5 \text{ g L}^{-1}$  peptone,  $50 \text{ g L}^{-1}$  agar-agar, and  $1.5 \text{ g L}^{-1}$  beef extract. For 200 ml of nutrient medium 10 g agar-agar, 1 g NaCl, 1 g peptone, 0.3 g beef extract were dissolved in 200 ml double distilled water. The solution was autoclaved at  $121^\circ\text{C}$  for 20 min. Before reaching  $45^\circ\text{C}$ , the solution was poured into petri dishes and left for solidification in a UV laminar chamber. After solidification, the nutrient media was inoculated with bacterial culture and incubated for 48 h for bacterial growth. These petri dishes containing bacterial colonies were subjected to antimicrobial test studies (fig 4). 2 mg of Penicillin was dissolved in 5 ml of water; in this solution two 6 mm What man paper discs were immersed. Similarly two more discs were immersed in  $\text{TiO}_2\text{-Ag}$  nanocomposites dissolved in 5 ml of ethanol. After some time (around 7 min) the discs were dried naturally, and then placed in petri dishes containing bacterial colonies. These petri dishes were incubated at  $38^\circ\text{C}$  for 48 h. The antimicrobial activity was assessed by observing and measuring inhibition zones formed around the discs after incubation period.

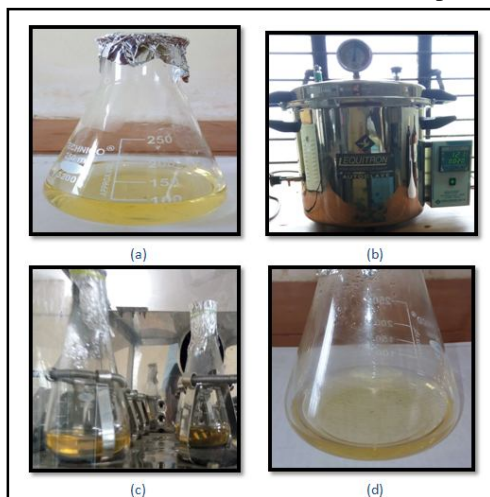


Fig 3: (a) LB medium before autoclaving, (b) Autoclaving of LB medium, (c) After bacterial inoculation placed in an orbital shaker for 8h (d) After orbital shaker LB medium showing bacterial growth.

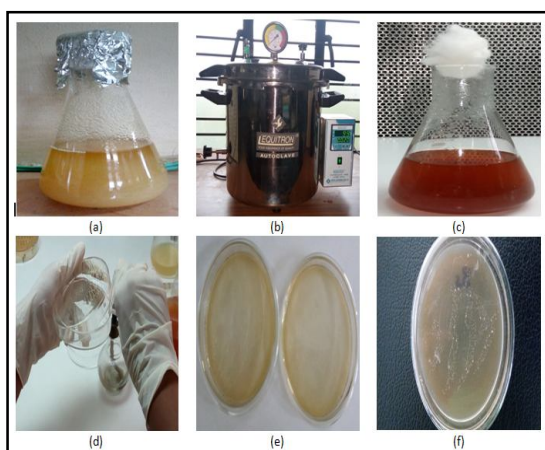


Fig 4: (a) Nutrient medium before autoclaving, (b) Autoclaving of nutrient medium, (c) Nutrient medium after autoclaving, (d) Nutrient medium poured into Petri dishes, (e) Solidification of the nutrient medium after pouring, and (f) Bacterial colonies after 48 h of inoculation.

### III. RESULTS & DISCUSSIONS

#### A. XRD Analysis

The crystalline nature of synthesized nanocomposites was identified using X-ray diffractogram (XRD) pattern (fig. 5). The diffractogram pattern was indexed properly for all crystalline peaks and compared with JCPDS data file. Figure (5) shows the major peaks at  $2\theta$  values of  $25.36^\circ$ ,  $38.01^\circ$ ,  $47.98^\circ$ ,  $54.76^\circ$  which corresponds to the planes of (101), (004), (200), and (211) of tetragonal anatase  $\text{TiO}_2$  (JCPDS Card No. 21-1272) [7],[8]. The peaks at  $2\theta$  values of  $37.26^\circ$ ,  $44.76^\circ$ ,  $64.85^\circ$  correspond to the planes of (111), (200), and (220) of face-centered cubic Ag (JCPDS card no. 04-0783) [9],[10]. The average crystallite size of  $\text{TiO}_2$ -Ag nanocomposites was estimated to be  $7.06 \pm 3.2$  nm from Debye-Scherrer formula. The composite showed crystalline phase peaks of both  $\text{TiO}_2$  & Ag and hence confirms the composite formation.

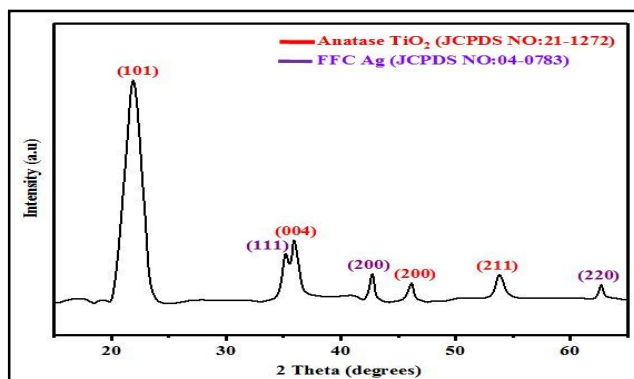


Fig 5: XRD analyses of  $\text{TiO}_2$ -Ag nanocomposites

#### B. FTIR Spectroscopy

FTIR analysis was performed to identify functional groups on surface of nanomaterials (fig6). The broad peak at  $500\text{-}900\text{ cm}^{-1}$  corresponds to the Ti-O-Ti stretching vibrations, whereas the peak at  $1384\text{ cm}^{-1}$  indicates vibrations of the Ti Ligands, peak at  $1633\text{ cm}^{-1}$  O-H bonds in water, peak at  $3450\text{ cm}^{-1}$  stretching vibrations of hydroxyl groups of Ti-OH on the surface. Further, the peak at  $3884\text{ cm}^{-1}$  attributes to the -OH stretching vibrations of alcohol [11], [12].

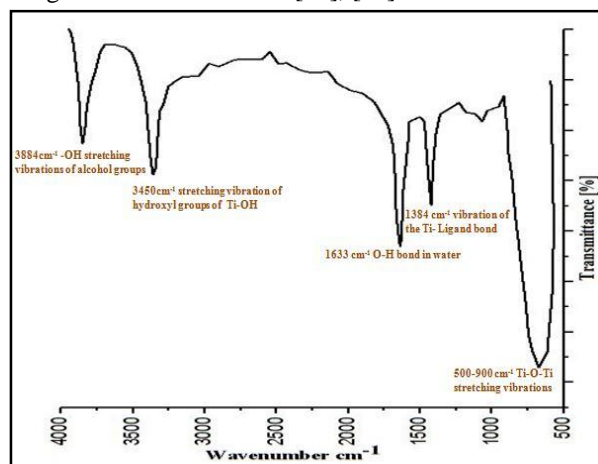
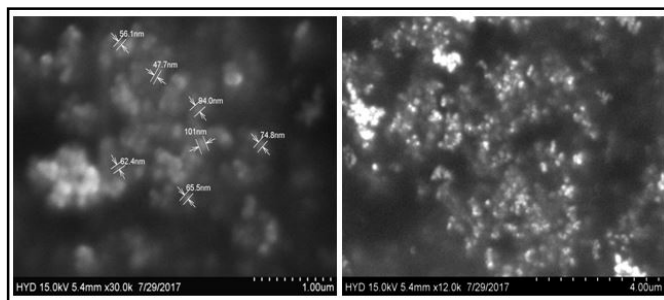


Fig 6: FTIR analyses of  $\text{TiO}_2$ -Ag nanocomposite

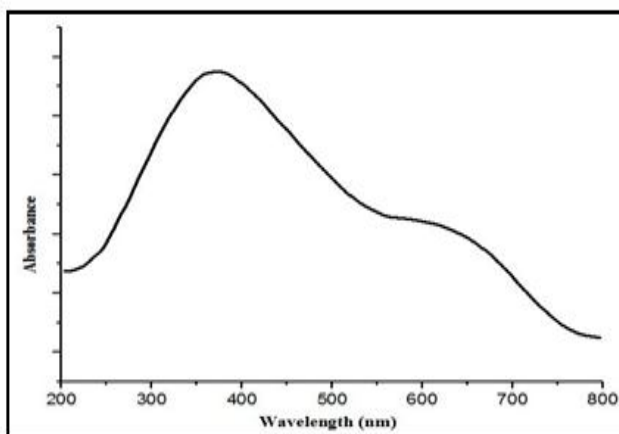
#### C. SEM Analysis

The particle size, shape and surface morphology characteristics of synthesized samples can be obtained from scanning electron microscopy (SEM). Figure (7) shows the image of  $\text{TiO}_2$ -Ag nanocomposites within the magnification of  $1\text{ }\mu\text{m}$ . The  $\text{TiO}_2$ -Ag nanocomposites contain irregularly shaped particles which may be due to the aggregation of tiny crystals [5]. The particle size ranged from  $47.7$  to  $101\text{ nm}$ .


Fig 7: SEM analyses of TiO<sub>2</sub>-Ag nanocomposites

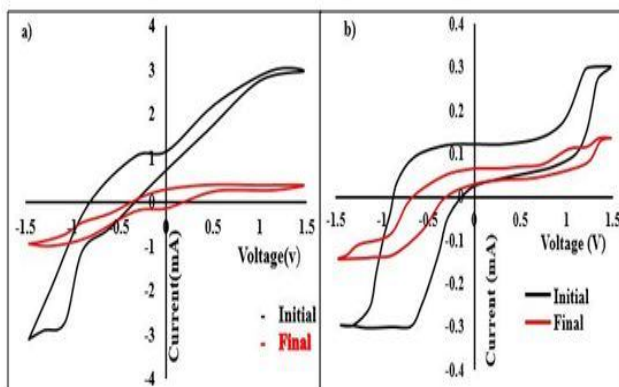
#### D. UV-Vis Spectroscopy

UV-Vis spectrometer was used to measure the optical properties of nanoparticles. Figure (8) shows the UV-Visible spectrum of TiO<sub>2</sub>-Ag nanocomposites. Because of surface plasmon resonance TiO<sub>2</sub> nanoparticles have an absorbance at 380 nm and Ag nanoparticles at 603 nm. Both nanoparticles exhibit absorbance in the visible region. Hence this nanocomposite works in the visible region for antimicrobial studies [11].


Fig 8: UV-Vis spectrum analyses of TiO<sub>2</sub>-Ag nanocomposites

#### E. Dye Degradation

1) *Cyclic Voltammetric Studies:* Cyclic Voltammetry (CV) studies are used to determine the rate of dye degradation before and after electrochemical treatment. During forward scan of the cyclic voltammetry, anodic (Oxidation) peaks were observed at positive currents; in reverse scan cathodic (Reduction) peaks were observed at negative currents. Oxidation peak indicates the presence of complex/reactive groups in dye molecules. The high peak current in initial CV indicates the presence of more number of reactive groups in dye molecules that leads to the higher oxidation of groups. To degrade these complex/reactive groups in crystal violet dye, potential (V) is applied on electrode that was chosen based on cyclic voltammetry (CV) curves.


Fig 9: Voltammetric plots of crystal violet for dye degradation using (a) Bare Ti and (b) TiO<sub>2</sub>-Ag@Ti.

Cyclic voltammetry plots correspond to working electrode of bare Ti and modified electrode of Ti/TiO<sub>2</sub>-Ag are shown in fig9(a and b). Dye degradation experiments were performed using Potentiostat. Potentiostat consists of three electrodes namely; the working electrode (anode) is fabricated electrodes, reference electrode is Ag/AgCl, and counter electrode (cathode) is Platinum mesh. Crystal violet dye (55 mg/L) dissolved in 1M Na<sub>2</sub>SO<sub>4</sub> solution is the electrolyte. The extent of dye degradation was explained with CV plots before (at t=0) and after electrolysis treatment (at t>0) of crystal violet dye. The cathodic and anodic peak currents originated are observed to have reduced with the time of electrolysis. The decrease in peak currents of crystal violet certified to the oxidation of dye. Moreover, this information is very much important to assess the achievability of the electrochemical process for the degradation of crystal violet dye. The oxidation current in CV curve represents the oxidation of dye molecules. The best performance of prepared electrodes was shown by the Ti/TiO<sub>2</sub>-Ag nanocomposites compared to bare TiO<sub>2</sub>. The TiO<sub>2</sub>-Ag nanocomposites contribute a free path to electrons in enhancing the conductivity leading to quick oxygen evolution [13].

2) *UV-Vis Spectroscopy Analysis*: During the electrochemical decolourization process, the change in absorbance characteristics of the crystal violet dye was observed at different intervals of time using UV-Vis absorption spectra analysis. The spectrum of crystal violet dye shows a maximum absorbance at 576 nm in the visible region representing the presence of methyl groups which is accountable for strong color of crystal violet dye. The absorption spectra were generated for various samples collected at different times during electrochemical treatment [14]. Decolourization curve of bare Ti is shown in figure (10 a), and that of TiO<sub>2</sub>-Ag@Ti shown in figure (10b). Complete decolorization was achieved within 14 hours for Ti, and 8 hours for TiO<sub>2</sub>-Ag. TiO<sub>2</sub>-Ag nanocomposites showed a faster rate of decolorization as Ag nanoparticles have dispersity on the particles of the TiO<sub>2</sub> support which reduces the band gap of TiO<sub>2</sub> nanoparticles and hence facilitates free electron path [15]. Further, the large surface area of the nanocomposites accommodates a number of dye molecules to interact with oxidants. Therefore, TiO<sub>2</sub>-Ag nanocomposites increase the catalytic performance and increase the rate of crystal violet dye degradation.

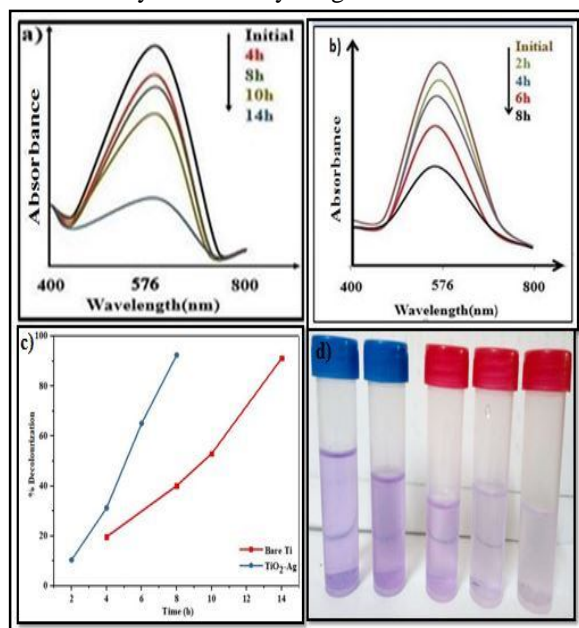


Fig 10: Absorbance V/s wavelength (nm) for (a) Bare Ti, and (b)TiO<sub>2</sub>-Ag@Ti (c) % Decolourisation V/s time for Bare Ti and Ti/TiO<sub>2</sub>-Ag (d) crystal violet dye at different intervals.

#### F. Antimicrobial Activity

The antimicrobial property of TiO<sub>2</sub>-Ag nanocomposites was evaluated against gram-positive bacteria *Bacillus subtilis* (*B. subtilis*) and gram-negative bacteria *Escherichia coli* (*E. coli*) by using disc diffusion method. In this method inhibitory activity was measured based on the diameter of the clear inhibition zone. If there was no surrounding clear zone, it was assumed that there was no inhibitory zone [16]. The antimicrobial activity was assessed by observing and measuring inhibition zones formed around the discs after incubation period. The zone of inhibition around Penicillin and TiO<sub>2</sub>-Ag nanocomposite discs are shown in fig (11) while numerical value of the diameter of inhibition zones was compiled in fig (12).

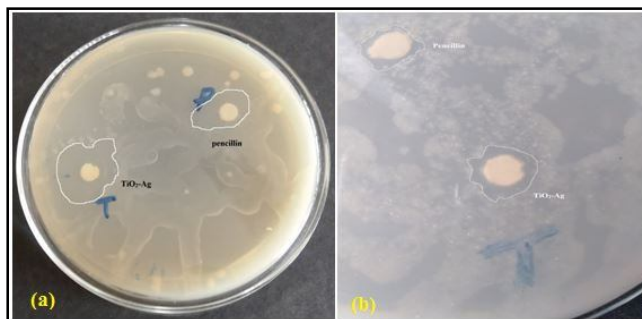


Fig 11: Antimicrobial activity assay of penicillin and  $\text{TiO}_2\text{-Ag}$  nanocomposites on (a) *B. subtilis* and (b) *E. coli*.

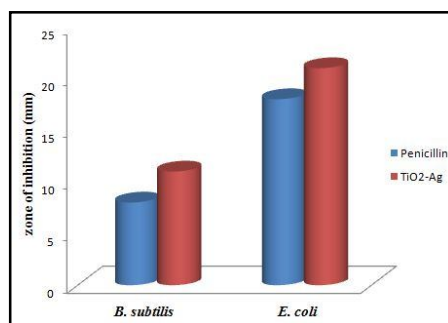


Fig 12: Graphical representation of antimicrobial activity

It is noticed from the results that the diameter of inhibition zone is higher for gram negative bacteria *Escherichia coli* (21 mm) than that of gram positive bacteria *Bacillus subtilis* (13 mm). This may be due to the difference in the compositions of the cell membranes [17].

#### IV. CONCLUSIONS

The  $\text{TiO}_2\text{-Ag}$  nanocomposites were successfully synthesized via green route from Aloe vera plant extract. The green synthesized nanocomposites have average particle size ranging from 47.4 to 101 nm. Average crystallite size was calculated as  $7.06 \pm 3.2$  nm. Surface plasmon resonance of the  $\text{TiO}_2$  and Ag nanoparticles was at 380 nm and 603 nm respectively. Ti functional groups were identified with the help of FTIR. In electrochemical dye degradation  $\text{TiO}_2\text{-Ag}$  nanocomposites enhanced the electrocatalytic performance compared to bare Ti. The antimicrobial studies showed that nanocomposites inhibited the gram-negative bacteria (*Escherichia coli*) more than that of gram-positive bacteria (*Bacillus subtilis*).

#### V. ACKNOWLEDGMENT

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#### REFERENCES

- [1] Khan, Mohammad Ehtisham, Mohammad Mansoob Khan, and Moo Hwan Cho. "Green synthesis, photocatalytic and photoelectrochemical performance of an Au-Graphene nanocomposite." *RSC Advances* 5.34 (2015): 26897-26904.
- [2] Shabani, Ali, et al. "Photo-catalyst  $\text{Fe}_3\text{O}_4$ ." *Journal of Materials Science: Materials in Electronics* 27.8 (2016): 8661-8669.
- [3] Prabhu, Sukumaran, and Eldho K. Poullose. "Silver nanoparticles: mechanism of antimicrobial action, synthesis, medical applications, and toxicity effects." *International Nano Letters* 2.1 (2012): 32.
- [4] Yang, Yucheng, et al. "Polypyrrole-decorated Ag- $\text{TiO}_2$  nanofibers exhibiting enhanced photocatalytic activity under visible-light illumination." *ACS applied materials & interfaces* 5.13 (2013): 6201-6207.



- [5] Jegadeeswaran, P., et al. "A novel green technology: Synthesis and characterization of Ag/TiO<sub>2</sub>nanocomposites using Padinatetrahromatica (seaweed) extract." Materials Letters 166 (2016): 137-139.
- [6] Jović, Milica, et al. "Study of the electrochemical oxidation of reactive textile dyes using platinum electrode." Int. J. Electrochem. Sci 8.1 (2013): 168-183.
- [7] Thesis: Surya PrakashGhosh, National Institute of Technology, Rourkela, Rourkela-769008,Orissa, India, (2012).
- [8] Habib, MdAhsan, et al. "Synthesis and characterization of ZnO-TiO<sub>2</sub>nanocomposites and their application as photocatalysts." International nano letters 3.1 (2013): 5.
- [9] Xu, Jian, et al. "Effect of silver ions on the structure of ZnO and photocatalytic performance of Ag/ZnO composites." Applied Surface Science 255.5 (2008): 1996-1999.
- [10] Zuas, Oman, NuryatiniHamim, and YuliantiSampora. "Bio-synthesis of silver nanoparticles using water extract of Myrmecodiapendan (SarangSemut plant)." Materials Letters 123 (2014): 156-159.
- [11] Sher Shah, MdSelimArif, et al. "Green synthesis of biphasic TiO<sub>2</sub>-reduced graphene oxide nanocomposites with highly enhanced photocatalytic activity." ACS applied materials & interfaces 4.8 (2012): 3893-3901.
- [12] Zeng, Tao, et al. "A novel Fe<sub>3</sub>O<sub>4</sub>-graphene-Au multifunctional nanocomposite: green synthesis and catalytic application." Journal of Materials Chemistry 22.35 (2012): 18658-18663.
- [13] Sharma, Rajni, et al. "ZnO anchored graphene hydrophobic nanocomposite-based bulk heterojunction solar cells showing enhanced short-circuit current." Journal of Materials Chemistry C2.38 (2014): 8142-8151
- [14] Mogal, Sajid I., et al. "Single-step synthesis of silver-doped titanium dioxide: influence of silver on structural, textural, and photocatalytic properties." Industrial & Engineering Chemistry Research 53.14 (2014): 5749-5758.
- [15] MilicaJović, DaliborStanković, DraganManojlović, Ivan Anđelković, AnđelijaMilić, Biljana Dojčinović1, GoranRoglić,Int. J. Electrochem. Sci., 8, 168 –183, (2013).
- [16] Awad, Manal A., et al. "Green synthesis, characterization, and antibacterial activity of silver/polystyrene nanocomposite." Journal of Nanomaterials 2015 (2015): 5.
- [17] Gopinath, V., et al. "Biosynthesis of silver nanoparticles from Tribulusterrestris and its antimicrobial activity: a novel biological approach." Colloids and Surfaces B: Biointerfaces96 (2012): 69-74.



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